

Markhor Junior Contest ®

Aspirant Level

(Class IX & X, O/A Level's) Science 2024

There are 25 questions, and each question carries two marks.

1. **Electromagnetic waves are classified according to their:**
 - a) Amplitude
 - b) Frequency
 - c) **Both A and B**
 - d) Neither A nor B

2. **The uncertainty principle states that it is impossible to know precisely:**
 - a) **The position and momentum of a particle simultaneously**
 - b) The energy and time of a particle simultaneously
 - c) The charge and mass of a particle simultaneously
 - d) The location and velocity of a star

3. **In a transistor, the flow of current between collector and emitter is controlled by the voltage applied to the:**
 - a) **Base**
 - b) Collector
 - c) Emitter
 - d) Gate

4. **The process by which rocks are formed from the accumulation of sediments is called:**
 - a) Weathering
 - b) Erosion
 - c) **Sedimentation**
 - d) Metamorphism

5. **Which of the following molecules is NOT a macromolecule?**
 - a) Protein
 - b) **Glucose**
 - c) DNA
 - d) Lipid

6. **The process by which plants convert light energy into chemical energy is called:**
 - a) Cellular respiration
 - b) **Photosynthesis**
 - c) Mitosis
 - d) Meiosis

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7. Which of the following is NOT a theory of evolution?
- a) **Lamarckism**
 - b) Natural selection
 - c) Gradualism
 - d) Punctuated equilibrium
8. The noble gases are unreactive because:
- a) **They have a full valence shell**
 - b) They have a high melting point
 - c) They are all metals
 - d) They are all radioactive
9. Which of the following is an example of an endothermic reaction?
- a) Combustion of methane
 - b) Dissolving salt in water
 - c) Photosynthesis
 - d) Neutralization reaction
10. The pH of a solution is a measure of the concentration of:
- a) **Hydrogen ions (H⁺)**
 - b) Hydroxide ions (OH⁻)
 - c) Salt ions
 - d) Water molecules
11. Which of the following is NOT a fundamental force in physics?
- a) Strong nuclear force
 - b) Electromagnetic force
 - c) Gravitational force
 - d) **Vital force**
12. The work done on an object is equal to the product of:
- a) **Force and distance**
 - b) Mass and acceleration
 - c) Power and time
 - d) Momentum and velocity
13. Lichens, a symbiotic relationship between a fungus and an alga/cyanobacteria, are considered pioneer species because they:
- a) **Can tolerate harsh environments**
 - b) Reproduce asexually very rapidly.
 - c) Are keystone predators in their ecosystem.
 - d) Fix nitrogen from the atmosphere.

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14. Homeostasis is crucial for maintaining a stable internal environment. Which of the following is NOT directly involved in thermoregulation in humans?
- a) Sweating to release excess heat.
 - b) Shivering to generate heat.
 - c) **Dilation of blood vessels in the skin to lose heat.**
 - d) Increased metabolic rate to generate heat.
15. Fullerenes are fascinating carbon cages with unique properties. Buckminsterfullerene (C₆₀) is a good conductor of electricity. This unexpected behavior is likely due to:
- a) The presence of multiple double bonds in the structure.
 - b) **Delocalization of electrons within the cage**
 - c) High melting point of the molecule.
 - d) Spherical shape of the C₆₀ molecule.
16. Dark matter is a mysterious component of the universe with significant gravitational influence. However, we cannot directly observe it because:
- a) It is located far away in distant galaxies.
 - b) **It does not emit or absorb light.**
 - c) It is composed of very small particles.
 - d) It exists in a different dimension.
17. Quantum mechanics challenges our classical understanding of the world. According to the Heisenberg uncertainty principle, it is impossible to know precisely:
- a) **The position and momentum of a particle simultaneously.**
 - b) The mass and energy of a particle.
 - c) The charge and spin of a particle.
 - d) The location and velocity of a planet.
18. In operons, the lac operon in E. coli for example, what regulatory molecule can bind to the operator region and prevent transcription of the downstream genes?
- a) Activator protein
 - b) RNA polymerase
 - c) **Repressor protein**
 - d) mRNA
19. In nuclear magnetic resonance (NMR) spectroscopy, which factor MOST influences the chemical shift of a particular nucleus within a molecule?
- a) The number of neutrons in the nucleus
 - b) **The electronegativity of surrounding atoms**
 - c) The type of bond the nucleus participates in
 - d) The mass of the molecule

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20. According to the theory of relativity, what happens to the mass of an object as its velocity approaches the speed of light?
- a) It remains constant.
 - b) It decreases.
 - c) **It increases towards infinity.**
 - d) It becomes zero.
21. What is the main product formed during the breakdown of ATP (adenosine triphosphate) to ADP (adenosine diphosphate) and inorganic phosphate (Pi)?
- a) Glucose
 - b) Water
 - c) Carbon dioxide
 - d) **Energy (in the form of a high-energy phosphate bond)**
22. How can the degeneracy of an energy level in a molecule be lifted?
- a) By increasing temperature
 - b) By applying an electric field
 - c) By exposing the molecule to light
 - d) **All of the above**
23. What is the characteristic feature that distinguishes fermions from bosons?
- a) Fermions have half-integer spin, while bosons have integer spin.
 - b) **Fermions obey the Pauli exclusion principle, while bosons do not.**
 - c) Fermions are made of quarks, while bosons are fundamental particles.
 - d) Fermions are involved in the strong nuclear force, while bosons are not.
24. In a double-slit experiment, what effect does increasing the slit separation have on the observed fringe pattern?
- a) The fringes become narrower.
 - b) **The fringes become wider.**
 - c) The central maximum becomes brighter.
 - d) The overall intensity of the pattern decreases.
25. What is the primary mechanism by which enzymes increase the rate of a chemical reaction?
- a) **By providing an alternative reaction pathway with a lower activation energy**
 - b) By lowering the overall energy change of the reaction
 - c) By increasing the concentration of reactants
 - d) By changing the equilibrium constant of the reaction